DECLARATION

I, Dewan Shahidur Rahman, hereby declare that the thesis entitled

"Photophysical Properties of Molecular Probes near Nanostructured

Metallic Surfaces" has not been submitted either in whole or in part

previously to any other institution for the award of any degree or

qualification and does not contain any previously published material or

written by another person, except where due reference is made in the text.

Place: Silchar

Date: 16.05.2016

Dewan Shahidur Rahman

Dedicated to My Beloved Parent Mrs. Joynab Begom Mr. Dewan Khalilur Rahman

13057



ASSAM UNIVERSITY: SILCHAR

(A CENTRAL UNIVERSITY CONSTITUTED UNDER ACT XIII OF 1989)

14550

INTEGRATED PRE Ph.D. COURSE WORK EXAMINATION, 2011

MARK SHEET

The following are the marks obtained by

DEWAN SHAHIDUR RAHMAN

Son/daughter of

DEWAN KHALILUR RAHMAN

and Roll 011110

and JOYNAB BEGOM of Department of Chemistry

bearing Registration Number 24401081 of 2006-2007

No

00210047

appeared in

INTEGRATED PRE Ph.D. COURSE WORK Examination, held in May' 2011

Course/Paper No	Name/Title of Course /Paper	Full Marks	Pass Marks	Marks Obtained	Grade Points
IPP-501	Research Methodology-I	100	50	59	5.9
IPP-502	Research Methodology-II	100	50	58	5.8
IPP-503	Chemistry	100	. 50	75	7.5
IPP-504	Term Paper	100	50	70	7.0

H. Carrier H. Carrier	GRAND TOTAL	262
	Average Grade Point (10 point scale)	6.55 (A)
ESULT: QUALIFIED for both M.Phil & Ph.D. admission.		

Qualifying marks for:

i) Ph.D. Admission: 55% in each paper.

ii) M.Phil Admission: 50% in each paper.

CONTROLLER OF EXAMINATIONS

Acknowledgements

At first, I would like to express my sincere thanks and deep sense of gratitude to Dr. S. K. Ghosh, my supervisor, whose thoughtful suggestions and guidance always support me in carrying out this research work. Without his constant supervision, motivation and inspiration at every step, perhaps this work could not have been matured up to the present level.

I would like to express my gratitude to Prof. P. C. Pal, the Head, Department of Chemistry and Ex. Heads, Prof. N. V. S. Rao and Prof. C. R. Bhattacharjee for their constant support and help during the research work.

I would like to thank Dr. Sk. Jasimuddin for providing facilities for cyclic voltammetry measurements. I would like to offer my gratitude to Prof. S. B. Paul, Dr. P. Mondal, Dr. M. K. Paul, Dr. D. Sengupta, Dr. H. Acharjee, Dr. S. Choudhury, Dr. T. Sanjoy Singh and Dr. R. Panchadhayee for their moral support and help during the course of this research work. I would also like to remember Prof. M. R. Islam for providing his laboratory at the initial stage to start this work.

I would like to thank Dr. Nikhil R. Jana, IACS, Kolkata for providing fluorescence microscopic facilities; Dr. Achintya Singh, Bose Institute, Kolkata for Raman measurements; SAIF, NEHU, Shillong for providing electron microscopic facilities and Dr. Soumen Basu, Thapar University for characterization of some of the materials.

I would like to convey my thanks to Mr. P. R. Ramesh, Mr. S. Bhattacharjee, Mr. B. Nath, Mr. Jahangir Alom Barbhuiya, Mr. Jamil Alom Barbhuiya, Mr. L. Marh, Mr. Rajib Kurmi, Mr. Sanjib Bagdi and all other office staffs for their valuable help during my Ph. D. programme.

I am grateful to my labmates, Mahammed Ali, Hirak Chatterjee, Hasimur Rahaman and Sudip Pal and my younger brother, Dewan Azharul Islam, for their untiring support during my Ph. D. work. They were not only the people I worked with, but also my friends in need.

I express my sincere thanks to all other Research Scholar of the Department of Chemistry, especially, Zeaul Hoque Mazumder, Sutapa Chakraborty, Kaushik Barman,

Rupam Chakraborty, Nayan Roy, Avijit Dutta, Nirmalendu Das, for their active cooperation and friendship that made me enjoyable during this research work.

I acknowledge Department of Chemistry, Assam University, Silchar, and funding agency, DST, DBT and UGC, New Delhi for providing infrastructures and financial support.

At last but not the least, I acknowledge and express my heartiest regards to my father, Mr. Dewan Khalilur Rahman and mother, Mrs. Joynab Begom, for their constant affection, motivation, inspiration and help. Without them, I would not been the person what I am today. I would like to thank all my family members for their constant support, encouragements and eagerness throughout my career.

Dewan Shahidur Rahman

Preface

The aim of this work is to investigate the unique photophysical properties of fluorophores and semiconductor materials near the metal nanostructures of various sizes and shapes and their formation of metal-fluorophore hybrid assembly, metalsemiconductor assembly and their studies of various photophysical and thermodynamic properties. For this purpose, the present Ph. D thesis entitled "Photophysical Properties of Molecular Probes near Nanostructured Metallic Surfaces" has been chosen. This thesis includes the synthesis of gold and silver nanoparticles of various sizes and shapes (spherical, dogbone nanorods, multispikes shaped nanostar) and the formation of metalfluorophore hybrid assembly using various fluorophores and their photophysical properties based on Förster resonance energy transfer and electron transfer mechanism including the studies of relative contribution of electron transfer. Moreover, size selective metal nanoparticles-semiconductor assembly and their various physical properties have been studied using microscopic (transmission electron microscopy, scanning electron microscopy) and spectroscopic (UV-visible spectroscopy, steady state and time resolved fluorescence spectroscopy, Raman spectroscopy, energy dispersive X-ray analysis, Fourier transform infrared spectroscopy, X-ray diffraction, and selected area electron diffraction) techniques and their catalytic and other application has been tried to develop in the present work. The complete work of the thesis has been divided into seven chapters.

Chapter 1

This chapter includes the general introduction about the current literature related to the research work presented on plasmonic metal nanoparticles and fluorophore molecules near to the surface of the nanostructures.

Chapter 2

The chapter two presents a brief description about the chemicals and reagents used during the experiments and various techniques and analytical instruments used for the characterization of the synthesized materials and their hybrid-assemblies.

Chapter 3

In this chapter, the synthesis of gold nanoparticles with variable sizes and their aggregation phenomena and fluorescence enhancement as well as fluorescence quenching

have been described. This chapter consists of two sub-chapters: (i) synthesis of ultrasmall gold nanoparticles with size ranges of 1-5 nm and compared the fluorescence enhancement by following the spontaneous and induced aggregation and (ii) synthesis of gold nanoparticles with variable sizes (1-55 nm) and their quenching phenomena have been explained by using different models like FRET and NSET mechanism and their critical limits.

Chapter 4

This chapter describes the relative electron transfer in fluorescence quenching of pyrene derivatives with silver nanoparticles of variable sizes. Here, relative electron transfer contribution of the fluorophore molecules and the size effect of the Ag NPs have been explained in the metal-probe hybrid system.

Chapter 5

This chapter consists of anisotropic gold nanostructures with two sub-chapters. Sub-chapter (i) explains the synthesis of dogbone shaped gold nanorods with variable aspect ratio and their effect in quenching process. Sub-chapter (ii) describes the synthesis of multispiked gold nanostructures and their various properties like fluorescence quenching, laser heating shape transformation, excess surface energy at the tips and stability based on thermodynamics.

Chapter 6

This chapter highlighted about the metal-semiconductor assembly and their photophysical behavior with size and composition in two separated sub-chapters. Firstly, the photophysical and catalytic properties of the composites made of zinc oxide quantum dots and gold nanospheres with different sizes have been studied. Secondly, the assembly formations of silver and zinc oxide of different compositions and their photophysical, microscopic and spectroscopic bahaviour have been studied

Chapter 7

The important outcomes of the completed research are summarized and the future scopes of the work are mentioned in this chapter.

TABLE OF CONTENTS

	Page
List of Tables	xiii
List of Figures	
List of Schemes	xxi
List of Symbols	
List of Abbreviations	xxiv
Contents	Page
1. General Introduction	1
Small is Different: A Brief Introduction to Nanoparticles	2
Optical and Electronic Properties of Nanoparticles	3
Fluorescent Probes	4
Physical Parameters Affecting Nanoparticle-Fluoroprobe Interaction	5
Theoretical Background	6
Förster Resonance Energy Transfer	7
Gersten-Nitzan (G-N) Model	9
CPS-Kuhn Model	11
Nanometal Surface Energy Transfer (NSET)	12
Experimental Observation	14
Fluorescence Quenching	14
Fluorescence Enhancement	16
Simultaneous Quenching and Enhancement	17
Photophysical Processes	18
Intermolecular and Intramolecular Interactions	19
Electron Transfer	19
Energy Transfer	20
Application of Metal-Probe Hybrid Assembly	21
References	22
2. Experimental	26

Experimental		27
	Reagents	27
	Instruments	27
3. Fluorescence	e of Molecules near Aggregated and Isolated Gold Nanoparticles	30
Fluorescer	nce Enhancement of Aggregated Gold Nanoclusters	31
	Introduction	31
	Experimental	34
Synthesis	and Characterization of PVP-Stabilized Gold	34
	Nanoparticles of Variable Sizes	
Synthesis	of Gold-Probe Assembly.	35
	Results and Discussion	35
Photophys	sical Properties of FITC.	35
Characteri	zation of the Gold Particles	36
Aggregation	on amongst the Ultrasmall Gold Nanoparticles	37
Fluorescei	nce Spectra of FITC in the Presence of Spontaneous	40
	and Induced-aggregated Gold Nanoparticles	
	Conclusion	49
Critical Po	oint of Relaying from Nanosurface Energy Transfer to Förster	50
Resor	nance Energy Transfer	
	Introduction	50
	Experimental	52
3.2.2.	1 Synthesis of Gold Nanoparticles of Variable Sizes	52
	Results and Discussion	55
	Conclusion	61
Reference	s	61
4. Relative Con	ntributions of Electron and Energy Transfer in Nanoparticle-	
Fluoroprobe S	ystems	66
Introduction	on	67
Experime	ntal	70
	Synthesis of Size-Selective Silver Nanoparticles	70
	Making Metal-Probe Hybrid Assembly	71

Results and Discussion	71
4. 4. Conclusion	86
4.5. References	87
5. Fluorescence of Molecules near Anisotropic Gold Nanostructu	ures 90
Aspect Ratio Dependent Fluorescence Quenching of Dogbo	ne-Shaped
Gold Nanorods	91
5. 1.1. Introduction	91
Experimental	93
Synthesis of Gold Dogbones of Variable Aspect Ratio	93
Preparation of Nanorod-Probe Hybrid Assembly	93
Results and Discussion	94
Conclusion	104
Excess Surface Energy at the Tips of Gold Nanospikes	105
Introduction	105
Experimental Section	107
Synthesis of Gold Nanospheres	107
Synthesis of Gold Nanospikes	107
Metal-Fluorophore Assembly Formation	108
Results and Discussion	108
Conclusion	126
References	126
6. Semiconductor Fluoroprobes near Metal Nanostructures	130
Manipulating Electron Transfer in ZnO-Au Nanohybrid	ls 131
Introduction	131
Experimental	133
Making Semiconductor-Metal Nanohybrids	133
Results and Discussion	135
Conclusion	152
Tunable Ultraviolet Photoluminescence of Hierarchical	ZnO–Ag 153
Assemblies	
Introduction	153

	Experimental Section	154
Synthe	sis of Zinc Oxide-Silver Superstructures using Different Concentrations of	of
Silver	Nitrate 154	
Results	s and Discussion	155
Conclu	ision	168
Re	eferences	168
7. Overall	Conclusions	174
Reprints		
Appendix1	L	178
	List of Tables	217
Table 3.1.	Synthetic Conditions and Characteristic Parameters of the Different	
	Sets of Gold Particles.	35
Table 3.2.	Synthetic Conditions and Characteristic Parameters of the Different	
	Sets of Gold Particles.	53
Table 3.3.	Synthetic Conditions and Characteristic Parameters of the Size-	
	Specific Gold Nanoparticles.	54
Table 4.1.	Synthetic Conditions for the Four Different Sets of Silver	
	Nanoparticles	71
Table 4.2.	Absorption and Emission Spectral Characteristics of the Molecular	
	Probes	74
Table 4.3.	Account of Related Parameters in Silver Nanoparticle-Induced	
	Fluorescence Quenching of Pyrene Moieties	84
Table 5.1.	Composition of the Growth Solutions for the Synthesis of 'Dogbone-	
	Shaped' Gold Nanorods	95
Table 6.1.	Synthetic Conditions and Characteristic Parameters of the Size-	
	Specific Gold Nanoparticles	134
Table 6.2.	Characteristic Physical Parameters as a Function of Gold Particle Size	
	in ZnO-Au Nanocomposites	150
Table 6.3.	Synthetic Conditions for Different Sets of ZnO-Ag Hybrid	
	Assemblies	155

List of Figures

Figure 1.1.	Energy levels in metal particles	3
Figure 1.2.	Excited-state deactivation processes in metal-fluorophore	
	nanohybrids	18
Figure 3.1.	UV-vis spectra of the as-synthesized colloidal gold of different	
	sizes	36
Figure 3.2.	TEM images of the PVP-stabilized gold nanoparticles of sets (a) C,	
	(b) D, (c) E and (d) G respectively	37
Figure 3.3.	(A) Evolution of surface plasmon band of gold nanoparticles (set	
	A) upon spontaneous aggregation at different time intervals and	
	(B) UV-vis spectra of seven different sizes of gold nanoparticles	
	upon spontaneous aggregation followed after 45 h for each set of	
	gold nanoparticles; (C) Evolution of surface plasmon band of gold	
	nanoparticles (set A) upon induced aggregation at different	
	concentrations of hydrazine and (D) UV-vis spectra of seven	
	different sizes of gold nanoparticles upon induced aggregation	
	measured after addition of 1.0 mL of hydrazine in each set of gold	
	nanoparticles	38
Figure 3.4.	TEM images of the gold nanoparticle (sets G) aggregates arising	
	due to (A) spontaneous and (B) induced aggregation	40
Figure 3.5.	Fluorescence spectra of FITC molecules in the presence of (a)	
	spontaneously and (b) induced-aggregated gold aggregates	
	containing different particle sizes in 1 min. of the preparation of	
	solution	42
Figure 3.6.	Relative fluorescence enhancement of FITC molecules in the	
	presence of (a) spontaneously and (b) induced aggregated gold	
	nanoparticles of different sizes in 1 min of the preparation of	
	solution	43
Figure 3.7.	Fluorescence spectra of FITC molecules in the presence of	
	different concentrations of (a) spontaneously and (b) induced	

	aggregated gold aggregates containing gold nanoparticles (sets A)	
	in 1 min. of the preparation of solution	45
Figure 3.8.	Relative fluorescence enhancement of FITC molecules in the	
	presence of (a) spontaneously and (b) induced aggregated gold	
	nanoparticles of different sizes in 1 min. of the preparation of	
	solution	46
Figure 3.9.	Fluorescence spectra of FITC molecules in the presence of (a)	
	spontaneously and (b) induced aggregated gold aggregates	
	containing gold nanoparticles (sets A) at different time intervals	47
Figure 3.10.	(A) Fluorescence intensity of FITC molecules at different time	
	intervals in the presence of (a) spontaneously and (b) induced	
	aggregated gold aggregates containing gold nanoparticles (sets A)	
	and (B) first order analysis of the curve in the presence of	
	spontaneously gold nanoparticles	47
Figure 3.11.	Absorption and emission spectra of FITC molecules in aqueous	
	solution. Inset shows the molecular structure of dye molecules	54
Figure 3.12.	Absorption spectra of gold nanoparticles of variable sizes of 1 to	
	55 nm	55
Figure 3.13.	TEM images of gold nanoparticles of variable sizes of set (a) A, (b)	
	B, (c) E, (d) F, (e) I, (f) L, (g) M, (h) N and (i) O respectively	56
Figure 3.14.	Surface area of the gold nanoparticles as a function of particle	
	diameter at constant mass	57
Figure 3.15.	Fluorescence quenching of FITC molecules in the presence of gold	
	nanoparticles of variable sizes	58
Figure 3.16.	Quenching efficiency of the gold nanoparticles of FITC	
	fluorescence as a function of particle diameter at constant mass	60
Figure 4.1.	(A, B, C, D) Normalized UV-visible spectra of silver nanoparticles	
	(50 μM) of variable sizes of sets A, B, C and D, respectively. Inset	
	shows the digital photograph of the corresponding silver	
	particles	71
Figure 4.2	(a h c d) Transmission electron micrographs of silver	

	nanoparticles of sets A, B, C and D, respectively; (e) high	
	resolution transmission electron micrograph and (f) selected area	
	electron diffraction pattern of set D silver particles	73
Figure 4.3.	Absorption (25, 25 & 25 μ M) and emission (0.3, 0.04 & 0.03 μ M)	
	spectra of pyrene, aminopyrene and aminomethylpyrene. Inset	
	shows the molecular structures of the fluorescent probes	74
Figure 4.4.	Fluorescence spectra of pyrene (0.3 µM), aminopyrene (0.04 µM)	
	and aminomethylpyrene (0.03 µM) in the presence of variable	
	concentrations of 6 nm silver nanoparticles	75
Figure 4.5.	Stern-Volmer plots showing the quenching efficiency of pyrene	
	(0.3 μ M), aminopyrene (0.04 μ M) and aminomethylpyrene (0.03	
	μM) in the presence of (A) 2, (B) 6, (C) 9 and (D) 21 nm silver	
	nanoparticles	76
Figure 4.6.	(A, B, C) Change in the surface plasmon band of 6 nm silver	
	nanoparticles (50 μ M) upon addition of pyrene (0.4 – 20 μ M),	
	aminopyrene (0.4 – 20 μ M), and aminomethylpyrene (0.4 – 20	
	μM), respectively and (D) relative dampening of absorbance of the	
	silver nanoparticles upon addition of fluoroprobes in different	
	concentrations. Inset in profile B shows the transmission electron	
	micrographs of the silver particles (set B) (a) before and (b) after	
	addition of aminopyrene molecules.	77
Figure 4.7.	Cyclic voltammograms of the pure fluoroprobes (4.5 μM), silver	
	nanoparticles (11.9 μ M) and silver (11.9 μ M)-fluoroprobe (4.5	
	μ M) hybrid assemblies at pH \sim 7.0	80
Figure 4.8.	Overlap spectra of the emission spectrum of pyrene (0.3 μ M),	
	aminopyrene (0.04 μ M), and aminomethylpyrene (0.03 μ M) with	
	the absorption spectrum of 6 nm silver nanoparticles (50 μ M)	82
Figure 4.9.	Overlap spectra of the emission spectrum of pyrene (0.3 μ M),	
	aminopyrene (0.04 μ M) and aminomethylpyrene (0.03 μ M) with 2,	
	6, 9 and 21 nm silver nanoparticles (50 μM)	83
Figure 4 10	Differential quenching of pyrene mojeties as a function of particle	

	size of silver nanostructures	85
Figure 5.1.	Absorption (0.16 mM) and emission spectrum (0.3 μ M) of Sudan	
	Red 7B in aqueous solution. Molecular structure of the dye is	
	shown in the inset	95
Figure 5.2.	Absorption spectra of gold dogbones (0.25 mM) of variable aspect	
	ratio. Inset shows TEM images of the (a) end and (b) side view	
	corresponding to transverse and longitudinal plasmon resonances	
	of the dogbones.	96
Figure 5.3.	TEM images of representative gold dogbones of variable aspect	
	ratio: (a) SP, (b) DNR 1, (c) DNR 3 and (d) DNR 6	97
Figure 5.4.	Emission spectra of SR7B (0.3 μ M) in the presence of GDNRs (20	
	μM) of variable aspect ratio	98
Figure 5.5.	Profile showing the relative intensity of SR7B as a function of	
	aspect ratio of GDNRs. Inset provides a schematic presentation	
	showing aspect ratio gradually increases from spheres to larger	
	dogbones.	99
Figure 5.6.	(A) Surface plasmon band (0.25 mM) of DNR4 upon successive	
	addition of SR7B. (Top) Inset shows the profile showing	
	dependence of $1/(-\Delta A)$ as a function of $1/[SR7B]$. (Bottom) Inset	
	shows a schematic presentation of the attachment of the dye	
	molecules on the surfactant-stabilized of GDNRs. (B) Overlap	
	between the absorption spectra of GDNRs of variable aspect ratio	
	(0.25 mM) and emission spectrum of SR7B molecules (0.3 μ M)	
	indicating efficient energy transfer	101
Figure 5.7.	Stern-Volmer plot of relative intensity as a function of	
	concentration of GDNRs of three different aspect ratio	103
Figure 5.8.	Absorption and emission spectra of alizarin red in aqueous	
	medium. Inset shows the molecular structure of the dye molecule.	108
Figure 5.9.	Absorption spectra of CTAB-stabilized gold nanospheres and	
	nanospikes. Inset shows the transmission electron microscopic	
	images(inset shows the corresponding diameter histogram) of gold	

	(a) nanospheres and (b) nanospikes	109
Figure 5.10.	Emission spectra of alizarin red molecules (33 μM) in presence of	
	different concentrations of gold (A) nanospheres and (B)	
	nanospikes in aqueous medium; (C) Relative efficiency of alizarin	
	red molecules in the presence of gold nanospheres and nanospikes	
	and (D) Stern-Volmer plot showing the quenching efficiency of	
	alizarin red as a function of concentration of gold nanospheres and	
	nanospikes	111
Figure 5.11.	(A) Absorption spectral changes of gold nanospikes (0.25 mM)	
	upon successive addition of alizarin red. Inset shows the	
	corresponding changes of the gold nanospheres upon addition of	
	the dye molecules. (B) Overlap between the emission spectrum of	
	alizarin red (33 µM) with absorption spectra of gold (0.25 mM)	
	nanospheres and nanospikes	114
Figure 5.12.	Temporal changes in the absorption spectral features of gold	
	nanospikes upon NIR laser irradiation. Inset shows the	
	transmission electron micrographs during the photoirradiation of	
	gold nanospikes to nanospheres at different stages of irradiation	116
Figure 5.13.	Emission spectra of alizarin red (33 µM) in the presence of	
	different morphology of gold nanostructures evolved during the	
	photoirradiation of gold nanospikes to nanospheres. Inset shows a	
	schematic presentation of the three-dimensional geometrical	
	modeling of the illustration of the particle shape evolved during	
	laser irradiation	117
Figure 5.14.	Exponential growth of the protruded arm: (A) two-dimensional	
	geometrical model and (B) corresponding three-dimensionally	
	rendered structure	118
Figure 5.15.	(a) Hollow bottom protrusion on a spherical core and (b) simplified	
	2D projection of a protruded cone onto a spherical core	119
Figure 5.16.	Plot showing relationship between surface energy (E), solvent	
	interaction energy parameter (k) and radius of curvature of the tip	

	(<i>r</i>)	122
Figure 5.17.	Modeling of the tip of a spike on the basis of intersection with a	
	plane: (panel A) symmetric and (panel B) asymmetrical inclination	
	of the cone, (panel C) the evolution of the conic section describing	
	a general shape of the non-spherical cap	123
Figure 5.18.	Three-dimensional modeling of the tip of a spike: (A) theoretical	
	model, (B) transmission electron micrograph and (C) three-	
	dimensional meshgrid of the rendered model of tip	124
Figure 5.19.	Change in surface energy of a tip ellipsoid as a function of	
	perpendicular semi-axis	124
Figure 5.20.	Change in excess surface energy of the nanospikes as a function of	
	radius of the evolved nanospheres	125
Figure 6.1.	Absorption spectra of ZnO QDs (16.7 µM) in the absence and	
	presence of Au NPs (0.67 μ M) of five different sizes. Inset shows	
	the change in the surface plasmon band of Au NPs (set B) upon	
	interaction with the ZnO QDs.	135
Figure 6.2.	(a, b) Transmission electron micrographs, (c, d) high resolution	
	transmission electron micrographs and (e, f) selected area electron	
	diffraction pattern of ZnO and ZnO-Au, respectively	137
Figure 6.3.	(A) Fluorescence spectra of ZnO QDs (16.7 μM) in the absence	
	and presence of Au NPs (0.67 µM) of five different sizes. Inset	
	shows the quenching efficiency of the Au NPs as a function of	
	particle size. (B) Stern -Volmer plots of the quenching of ZnO	
	QDs (16.7 μ M) upon addition of five different sizes of Au NPs	138
Figure 6.4.	(A) Absorption spectral changes of ZnO QDs (16.7 μM) upon	
	successive addition of gold nanoparticles (set B) (4 – 40 μ M). Inset	
	shows the profile showing dependence of $1/(-\Delta A)$ as a function of	
	1/[Au NPs]. (B) Overlap between the UV emission spectrum of	
	ZnO QDs (16.7 μ M) with absorption spectra of gold nanoparticles	
	(40 μM) of five different sizes	140
Figure 6.5	Fourier transform infraredspectra of the as-prepared (a) $7nO$ ODs	

	and (b) ZnO–Au NCs	142
Figure 6.6.	Cyclic voltammograms of ZnO QDs (1.67 mM), Au NPs (0.067	
	mM), and ZnO-Au NCs (molar ratio 25:1) in 0.1 M KCl solution	
	in different potential windows	143
Figure 6.7.	X-ray diffraction pattern of the as-prepared (a) ZnO QDs and (b)	
	ZnO–Au NCs	144
Figure 6.8.	Enhanced multiphonon Raman scattering spectrum of ZnO QDs	
	and ZnO–Au NCs.	145
Figure 6.9.	(A) Absorption spectral changes during the degradation of Evans	
	blue (20 µM) in the presence of ZnO-Au NCs (containing gold	
	nanoparticles from set B) upon visible light irradiation. (B) Profile	
	showing the plot of $ln(A_0/A)$ as a function of time for	
	photocatalytic degradation of Evans blue in the presence of ZnO	
	QDs and ZnO-Au NCs containing gold nanoparticles of five	
	different sizes	148
Figure 6.10.	Profile showing the correlation between the apparent association	
	constants, Stern-Volmer quenching constants and catalytic rate	
	constants with the band gap as a function of particle diameter of	
	gold in ZnO-Au nanohybrids	151
Figure 6.11.	(a-e) Scanning electron micrographs of ZnO-Ag assemblies with	
	ZnO: Ag = $1.0 : 0.01, 1.0 : 0.013, 1.0 : 0.02, 1.0 : 0.04$ and $1.0 : 0.04$	
	0.1, respectively; (f) high resolution SEM image showing the	
	epitaxial growth of sets C, (g) fluorescence microscopic image (set	
	B) and (h) energy dispersive X-ray spectrum (set B)	156
Figure 6.12.	(a, b) Transmission electron micrographs of sets A and E,	
	respectively, (c) high resolution transmission electron micrograph	
	of set E and (d) selected area electron diffraction pattern of set E,	
	respectively corresponding to ZnO–Ag assemblies	158
Figure 6.13.	Diffuse reflectance spectra of pure ZnO and different sets of ZnO-	
	Ag assemblies	159
Figure 6.14.	Tauc's plot for the estimation of band gap energy of pure ZnO and	

	different sets of ZnO–Ag assemblies	160
Figure 6.15.	Fourier transform infrared spectrum of (a) pure ZnO and (b) ZnO-	
	Ag assemblies (set A)	161
Figure 6.16.	Cyclic voltammograms of different sets of ZnO-Ag assemblies	
	(0.01 M)	161
Figure 6.17.	X-ray diffraction patterns of pure ZnO and different sets of ZnO-	
	Ag assemblies	162
Figure 6.18.	Raman spectra of pure ZnO and different sets of ZnO-Ag	
	assemblies.	163
Figure 6.19.	Kinetic studies by measuring (A) absorption, (B) Fourier transform	
	infrared and (C) X-ray diffraction pattern showing the evolution of	
	ZnO-Ag assemblies of set E at different time intervals	165
Figure 6.20.	(A) Photoluminescence spectra (in the solid state) and (B)	
	deconvolution of the spectra of pure ZnO and different sets of	
	ZnO–Ag assemblies	167
	List of Schemes	
Scheme 3.1.	Molecular structure of FITC	36
Scheme 3.2.	Schematic presentation of the plausible binding modes of the dye	
	molecules to the (A) spontaneous and (B) induced-aggregated gold	
	nanoparticles	48
Scheme 4.1.	Schematic presentation of the possible deactivation pathways in	
	silver-fluoroprobe hybrid assemblies	. 85
Scheme 6.1.	Schematic presentation of miniaturising the band gap in ZnO-Au	
	nanohybrid assemblies containing Au NPs of five different sizes.	136
Scheme 6.2.	Schematic presentation of the photocatalytic degradation of Evans	
	blue at the ZnO-Au nanocomposites surfaces	149
		1 1/

Symbols

2D	two dimensional	3D	three dimensional
Å	angstrom	A_e	surface area of ellipsoidal tip-head
A_p	surface area of protrusion base	A_{cone}	surface area of prolong conical
A_{core}	surface area of spheroidal core		arm
c	velocity of light	°C	celsius
cm	centimeter	C	concentration
e	ellipticity of the tips of		distance of the emitter from the
	nanospikes		metal surface
E_{g}	band gap	eV	electron volt
E_{tip}	surface energy on tip	ε, εΑ	extinction co-efficient
E_{pc}	cathodic potentials	E_{pa}	anodic potentials
	energy transfer efficiency		Fermi energy level
ΔE_{P}	potential difference		real dielectric of the medium
()	real components of the	()	imaginary components of the
	dielectric function		dielectric function
f	packing fraction	$F_{\rm D}(\lambda)$	fluorescence intensity of donor
Φ_{dye}	quantum yield of the donor dye	fcc	face-centered cubic
ϕ_{f}	quantum yield		modified quantum yield
	radiative decay rate		nonradiative decay rate
h	hour	hν	photon energy
i_{pa}	anodic peak current	i_{pc}	cathodic peak current
	intensity of the exciting source	k	solvent interaction energy
	beam		parameter
$J(\lambda)$	overlap integral	$K_{\rm app}$	apparent association constant
K	kelvin		energy transfer rate
kDa	kilodalton	kV	kilovolt
K_S	static quenching constant	K_D	dynamic quenching constant
K_{SV}	Stern-Volmer constant	K_{cat}	catalytic rate constant
	Fermi wave-vector	λ_{ex}	excitation wavelength

λ_{max}	maximum wavelength	λ_{abs}	maximum absorption wavelength
λ_{em}	maximum emission wavelength	M	molar
mM	millimolar	mL	milliliter
meV	millielectron volt	min	minutes
N	numbers of nanoparticles per	nm	nanometer
	milliliter	nM	nanomolar
	Avogadro's number		refractive index
μM	micromolar	μA	microampere
	dipole moment		quantum yield of the donor
r	radius of atoms	R	average radii of the particles
R_{θ}	Förster distance		lifetime of the donor
	lifetime of the dye		excited state radiative lifetime of
Γ_0	excited state life time		the donor
	extinction of the particles	V_0	volume of the particles
ω	angular frequency		plasma frequency
	vibrational frequency of the	W	watt
	dipole		

Abbreviations

A	acceptor	AMPY	aminomethylpyrene
APY	aminopyrene	AR	alizarin red
Au NPs	gold nanoparticles	Ag NPs	silver nanoparticles
Ag-APY	silver-aminopyrene assembly	Ag-AMPY	silver-aminomethylpyrene
Ag-PY	silver-pyrene assembly		assembly
CTAB	cetyltrimethylammonium	CV	cyclic voltammetry
	bromide	D	donor
D	diameter	DNA	deoxyribonucleic acid
DNR	Dogbone-shaped nanorod	EB	Evans blue
FTIR	Fourier transform infrared	FITC	fluorescein isothiocyanate
	spectrum	FDTD	finite difference time
FRET	Förster resonance energy		domain
	transfer	GDNR	gold dogbone nanorod
HRTEM	high resolution transmission	LSPR	localized surface plasmon
	electron microscopy		resonances
NIR	near infrared	NPs	nanoparticles
NSET	nanosurface energy transfer	NCs	nanocomposites
PY	pyrene	PET	photoinduced electron
PVP	poly(<i>N</i> -vinyl-2-pyrrolidone)		transfer
PZC	point of zero charge	Q	quencher
QD	quantum dot	SP	sphere
SPR	surface plasmon resonance	SR7B	sudan red 7B
TEM	transmission electron	SAED	selected area electron
	microscopy		diffraction
TOAB	tetraoctylammonium bromide	UV	ultraviolet
UV-vis	ultraviolet visible	XRD	X-ray diffraction
UV-vis-	ultraviolet visible near	ZnO QD	zinc oxide quantum dot
NIR	infrared	ZnO-Au	zinc oxide-gold

ZnO-Au NCs zinc oxide-gold nanocomposites